Chapter 10 Study Guide-ANSWER THESE ON A SEPARATE SHEET OF PAPER AND TURN THEM IN THURSDAY

1. What does a superscript tell you and where is it located?
2. What does a subscript look like and where is it located?
3. What are valence electrons and why are they important?
4. What is the maximum number of valence electrons that can fit in an electron dot diagram?
5. When naming ionic compounds, which ion is listed last?
6. When naming compounds, what do you change the ending of the second ion to? What determines what ending that you choose?
7. What is a double bond?
8. How many valence electrons do noble gases have? Are they stable? Why or why not?
9. When an atom loses an electron, what happens? When it gains an electron?
10. What is an ionic bond?
11. What are ions that are made of more than one atom?
12. Do ionic compounds have a charge? Why or why not?
13. What is a covalent bond?
14. What is one reason that metals have the properties like luster, malleability, ductility, and conductivity?
15. What are polar and nonpolar bonds?
16. What is the pH and characteristics of an acid?
17. What is the pH and characteristics of a base?
18. What happens when an acid and base are mixed?
19. What do the groups and periods of the periodic table tell you?
20. What does the atomic number tell you?
21. How do you determine the number of neutrons in an element?
22. Explain how to do a Bohr diagram.
23. Explain how to do an electron dot diagram.
24. What are an ion, isotope, and molecule?
25. How do you form a new element?

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